

**METHOD FOR INCREASING LIGHT OLEFIN
YIELD BY CONVERSION OF A HEAVY
HYDROCARBON FRACTION OF A PRODUCT TO LIGHT OLEFINS**

5 **Field of the Invention**

 The present invention is directed to a method for increasing the yield of light olefins during the conversion of oxygenates to olefins by converting heavy hydrocarbons in the product to light olefins either by (a) recycling the heavy hydrocarbons to the primary reactor after removing the light olefins, or
10 (b) converting the heavy hydrocarbons to light olefins in a separate auxiliary reactor.

Background of the Invention

 Light olefins (defined herein as "ethylene and propylene") serve as feeds for the production of numerous chemicals. Light olefins traditionally are
15 produced by petroleum cracking. Because of the limited supply and/or the high cost of petroleum sources, the cost of producing olefins from petroleum sources has increased steadily.

 Alternative feedstocks for the production of light olefins are oxygenates, such as alcohols, particularly methanol, dimethyl ether, and
20 ethanol. Alcohols may be produced by fermentation, or from synthesis gas derived from natural gas, petroleum liquids, carbonaceous materials, including coal, recycled plastics, municipal wastes, or any organic material. Because of the wide variety of sources, alcohol, alcohol derivatives, and other oxygenates have promise as an economical, non-petroleum source for
25 olefin production.

 Because light olefins are the most sought after products of such a reaction, a continuing need exists for new methods to increase the yield of light olefin products and reduce the yield of unwanted products, such as "heavy" hydrocarbons having molecular weights heavier than propane.

Summary of the Invention

The present invention provides a method for increasing light olefin yield during conversion of oxygenates to olefins. The method comprises: contacting an oxygenate feed in a primary reactor with a non-zeolitic molecular sieve catalyst under first conditions effective to produce a first product comprising light olefins; separating the first product into light olefins and a heavy hydrocarbon fraction; feeding the heavy hydrocarbon fraction either back to the primary reactor or to a separate auxiliary reactor; and, subjecting the heavy hydrocarbon fraction to second conditions effective to convert at least a portion of the heavy hydrocarbons to light olefins.

Detailed Description of the Invention

In the conversion of oxygenates to light olefins (defined herein as ethylene and propylene), it is desirable to maximize the yield of light olefins. The present invention maximizes the yield of light olefins by converting the "heavy hydrocarbon fraction" of the reaction product to olefins. The "heavy hydrocarbon fraction" is defined herein as the fraction containing hydrocarbons having a molecular weight greater than propane. The conversion of heavy hydrocarbons to light olefins is accomplished either by (a) returning all or a part of the heavy hydrocarbon fraction to the primary reactor, where the heavy hydrocarbons are converted to light olefins along with additional oxygenate feed, or (b) conveying the heavy hydrocarbon fraction to a separate auxiliary reactor where the heavy hydrocarbons are converted to light olefins.

Molecular sieve catalysts that are suitable for use in the primary reactor are non-zeolitic catalysts, which include, but are not necessarily limited to silicoaluminophosphates ("SAPO's"). SAPO's have a three-dimensional microporous crystal framework of PO_2^+ , AlO_2^- , and SiO_2 tetrahedral units. Preferred SAPO's for use in the primary reactor are "small" and "medium" pore SAPO's. "Small pore" molecular sieve catalysts are defined as catalysts with pores having a diameter of less than about 5.0

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3

Angstroms. "Medium pore" molecular sieve catalysts are defined as catalysts with pores having a diameter in the range of from about 5 to about 10 Angstroms.

Suitable SAPO's for use in the invention include, but are not necessarily limited to SAPO-11, SAPO-44, SAPO-34, SAPO-17, and SAPO-18. A preferred SAPO is SAPO-34, which may be synthesized according to US-A-4,440,871, incorporated herein by reference, and Zeolites, Vol. 17, pp. 512-522 (1996), incorporated herein by reference.

SAPO's with added substituents also may be useful in the present invention. These substituted SAPO's form a class of molecular sieves known as "MeAPSO's." Suitable substituents include, but are not necessarily limited to nickel, cobalt, strontium, barium, and calcium.

Any molecular sieve catalyst capable of converting hydrocarbons with 4 or more carbon atoms into light olefins may be used in an auxiliary reactor. Preferred molecular sieve catalysts for the auxiliary reactor are zeolites.

Structural types of zeolites that are suitable for use in the auxiliary reactor with varying levels of effectiveness include, but are not necessarily limited to AEI, AFT, APC, ATN, ATT, ATV, AWW, BIK, CAS, CHA, CHI, DAC, DDR, EDI, ERI, GOO, KFI, LEV, LOV, LTA, MON, PAU, PHI, RHO, ROG, and THO and substituted examples of these structural types, as described in W. M. Meier and D. H. Olsen, Atlas of Zeolite Structural Types (Butterworth Heineman-3rd ed. 1997), incorporated herein by reference. Structural types of medium pore molecular sieves useful in the present invention include, but are not necessarily limited to, MFI, MEL, MTW, EUO, MTT, HEU, FER, AFO, AEL, TON, and substituted examples of these structural types, as described in the Atlas of Zeolite Types, previously incorporated herein by reference. A preferred zeolite for the auxiliary reactor is ZSM-5.

The process for converting oxygenates to olefins employs an organic starting material (feedstock) preferably comprising "oxygenates." As used herein, the term "oxygenates" is defined to include, but is not necessarily

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4

limited to aliphatic alcohols, ethers, carbonyl compounds (aldehydes, ketones, carboxylic acids, carbonates, and the like), and also compounds containing hetero-atoms, such as, halides, mercaptans, sulfides, amines, and mixtures thereof. The aliphatic moiety preferably should contain in the range of from about 1 to about 10 carbon atoms and more preferably in the range of from about 1 to about 4 carbon atoms. Representative oxygenates include, but are not necessarily limited to, lower straight chain or branched aliphatic alcohols, their unsaturated counterparts, and their nitrogen, halogen and sulfur analogues. Examples of suitable compounds include, but are not necessarily limited to: methanol; ethanol; n-propanol; isopropanol; C₄ - C₁₀ alcohols; methyl ethyl ether; dimethyl ether; diethyl ether; di-isopropyl ether; methyl mercaptan; methyl sulfide; methyl amine; ethyl mercaptan; di-ethyl sulfide; di-ethyl amine; ethyl chloride; formaldehyde; di-methyl carbonate; di-methyl ketone; n-alkyl amines, n-alkyl halides, n-alkyl sulfides having n-alkyl groups of comprising the range of from about 3 to about 10 carbon atoms; and mixtures thereof. As used herein, the term "oxygenate" designates only the organic material used as the feed. The total charge of feed to the reaction zone may contain additional compounds such as diluents.

Preferably, the oxygenate feedstock should be fed to the primary reactor and contacted in the vapor phase in a reaction zone with the selected molecular sieve catalyst at effective process conditions so as to produce the desired olefins, i.e., an effective temperature, pressure, WHSV (Weight Hourly Space Velocity) and, optionally, an effective amount of diluent, correlated to produce olefins. Alternately, the process may be carried out in a liquid or a mixed vapor/liquid phase. When the process is carried out in the liquid phase or a mixed vapor/liquid phase, different conversions and selectivities of feedstock-to-product may result depending upon the catalyst and reaction conditions.

The temperature employed in the primary reaction zone may vary over a wide range depending, at least in part, on the selected catalyst. Although

not limited to a particular temperature, best results will be obtained if the process is conducted at temperatures in the range of from about 200°C to about 700°C, preferably in the range of from about 250°C to about 600°C, and most preferably in the range of from about 300°C to about 500°C. Lower temperatures generally result in lower rates of reaction, and the formation of the desired light olefin products may become markedly slow. However, at higher temperatures, the process may not form an optimum amount of light olefin products, and the coking rate may become too high.

When an auxiliary reactor is used, the temperature range may be optimized to more effectively convert C_4^+ hydrocarbons to light olefins. Although not limited to a particular temperature, best results will be obtained if the process is conducted at temperatures within the range of from about 300°C to about 750°C, preferably in the range of from about 450°C to about 700°C. Lower temperatures generally result in lower rates of reaction and the formation of desired lower olefins may become markedly slow. However, at higher temperatures, the process may not form an optimum amount of light olefin products and the coking rate may become too high.

Light olefin products will form in both the primary reaction zone and in the optional auxiliary reactor--although not necessarily in optimum amounts--at a wide range of pressures, including but not limited to autogeneous pressures and pressures in the range of from about 0.1 kPa to about 100 MPa. A preferred pressure is in the range of from about 6.9 kPa to about 34 MPa, most preferably in the range of from about 48 kPa to about 0.34 MPa. The foregoing pressures are exclusive of diluent, if any is present, and refer to the partial pressure of the feedstock as it relates to oxygenate compounds and/or mixtures thereof. Pressures outside of the stated ranges may be used and are not excluded from the scope of the invention. Lower and upper extremes of pressure may adversely affect selectivity, conversion, coking rate, and/or reaction rate; however, light olefins such as ethylene still may form.

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6

The process should be continued for a period of time sufficient to produce the desired olefin products. The reaction cycle time may vary from tenths of seconds to a number of hours. The reaction cycle time is largely determined by the reaction temperature, the pressure, the catalyst selected, the weight hourly space velocity, the phase (liquid or vapor), and the selected process design characteristics.

A wide range of weight hourly space velocities (WHSV), defined as weight feed per hour per weight of catalyst, for the feedstock will function in the present invention. The WHSV generally should be in the range of from about 0.01 hr^{-1} to about 5000 hr^{-1} , preferably in the range of from about 1.0 hr^{-1} to about 2000 hr^{-1} , and most preferably in the range of from about 20 hr^{-1} to about 1000 hr^{-1} . The catalyst may contain other materials which act as inerts, fillers, or binders; therefore, the WHSV is calculated on the feed rate of oxygenate and recycle hydrocarbons and weight of inert, filler, and binder free catalyst.

When an auxiliary reactor is used, a wide range of WHSV's may be used in the secondary conversion reactor. The WHSV generally should be in the range of from about 0.1 hr^{-1} to about 5000 hr^{-1} , preferably in the range of from about 1.0 hr^{-1} to about 1000 hr^{-1} , and most preferably in the range of from about 1.0 hr^{-1} to about 100 hr^{-1} . The catalyst may contain other materials which act as inerts, fillers, or binders; therefore, the WHSV is calculated on the basis of feed rate of hydrocarbon and weight of inert, filler, and binder free catalyst.

One or more diluents may be fed to the reaction zone with the oxygenates, such that the total feed mixture comprises diluent in a range of from about 1 mol % and about 99 mol %. Diluents which may be employed in the process include, but are not necessarily limited to, helium, argon, nitrogen, carbon monoxide, carbon dioxide, hydrogen, water, paraffins, other hydrocarbons (such as methane), aromatic compounds, and mixtures thereof. Preferred diluents are water and nitrogen.

It may be desirable to provide for C_4^+ product purge at some point in the cycle to provide a ready source of fuel gas for plant use, C_4^+ product for sale, or to limit the size of the recycle stream to the primary reactor.

A preferred embodiment for the primary reactor is a circulating fluid bed or riser reactor with continuous regeneration, similar to a modern fluid catalytic cracker. Fixed bed reactors are not practical for use in the primary reactor because oxygenate to olefin conversion is highly exothermic and requires several stages with intercoolers or other cooling devices. The reaction also results in a high pressure drop due to the production of low pressure, low density gas.

Preferred auxiliary reactors include fluid bed, fixed bed, or fired tube reactors. The C_4^+ cracking process is endothermic; therefore, the reactor should be adapted to supply heat as necessary to maintain a suitable reaction temperature. On-line or periodic regeneration of the catalyst may be provided by any suitable means known in the art.

The invention will be better understood with reference to the following examples which are intended to illustrate, but not to limit the present invention.

EXAMPLE I

A sample of 50.7 mg of SAPO-34 catalyst obtained from UOP, Des Plaines, Illinois, was calcined at 650°C for 16 hours in air, pelletized, and sieved to produce 30 x 60 US mesh sized particles. The catalyst particles were loaded into a 4 mm diameter quartz reactor tube without further treatment. The quartz reactor tube was loaded into an electrically heated zone in the carrier gas stream of an HP 5890 gas chromatograph. Pressure was maintained at 25 psig (273 kPa) for the experiment. Pulses of 1-butene, 1-butene + methanol and pure methanol were alternatingly passed through the catalyst bed, products were analyzed on-line using a Flame Ionization Detector. Feed consisted of alternating pulses of 1 μ l of methanol, 1 ml of 1-butene + 1 μ l methanol, or 1 ml of 1-butene, respectively. Conversions were calculated based on water and coke free hydrocarbon production in the methanol feed case, and on water, coke and C_4 free hydrocarbon production in both the 1-butene alone and 1-butene plus methanol cases. The following

were the results, in weight percent selectivity:

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Feed	Conversion	CH ₄	C ₂ ⁼	C ₃ ⁼	C ₄ ⁼	C ₅ ⁺
1-butene	34.23	0.10 0.29	10.57	66.29	N/A	21.74
MeOH	100.00	0.86	51.06	34.19	9.29	2.08
MeOH + 1-butene	84.32	1.09	45.72	41.46	N/A	2.67

The foregoing results establish that the reaction of butene over an oxygenate-to-olefin catalyst produces propylene, ethylene, and C₅⁺ hydrocarbons. Additionally, these tests establish that, if both methanol and 1-butene are fed over an oxygenate-to-olefins catalyst, results intermediate between the results for methanol and 1-butene alone are achieved.

EXAMPLE II

The procedures of Example I were repeated except that only 1-butene was used as the feedstock in the process, and two catalysts were used. The first catalyst was SAPO-34 obtained from UOP. The second catalyst was silicalite bound ZSM-5. The results were as follows, in weight percent selectivity:

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Catalyst	Conversion	CH ₄	C ₂ ⁼	C ₃ ⁼	C ₄ ⁼	C ₅ ⁺
SAPO-34	34.23	0.10 0.29	10.57	66.29	N/A	21.74
ZSM-5	73.10 73.45	0.21 0.34	22.31 18.43	53.03 41.34	N/A	19.91 66.27

The foregoing demonstrates that both SAPO-34 and ZSM-5 convert 1-butene to ethylene and propylene, ZSM-5 having a higher overall conversion rate.

Persons of ordinary skill in the art will recognize that many modifications may be made to the present invention without departing from

the spirit and scope of the present invention. The embodiment described herein is meant to be illustrative only and should not be taken as limiting the invention, which is defined in the following claims.

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